Work Sheet 1/6

Q1. What is meant by reversible and irreversible reactions? Give an example of each.

Q2. What is physical equilibrium? Give examples.

Q3. State and explain Henry’s law? Write its applications.

Q4. What are the general characteristics of physical equilibrium?

Q5. Mention the conditions when Henry’s law is applicable.

Q6. What are the characteristics of equilibrium state?

Q7. What are homogeneous and heterogeneous reversible reactions?

Q8. Which measurable property becomes constant in H2O(l) $⇌$ H2O(vap) at equilibrium state

Q9. Under what condition does a reversible physical equilibrium process become irreversible?

Q10. Give the conditions at which a chemical equilibrium can exists in an open vessel. Illustrate your answer with suitable example.

Q11. A liquid is in equilibrium with its vapour in a sealed vessel at a fixed temperature. The volume of container is suddenly doubled: (i) What is the initial effect of the change on vapour pressure? (ii) How do the rate of evaporation and condensation change initially? (iii) What is about the final vapour pressure when equilibrium is restored?